

## Esercitazioni Numeriche e di Laboratorio di Chimica di Base

### Soluzioni Esercizi Capitolo 2

2.1 R = 22.0 g

2.2 R = 9.260 g

2.3 R = 5.76 g

2.4 R = 71.45 g

2.5 R = 83.4 g

2.6 R = 95.0 %

2.7 R = 15.1 g

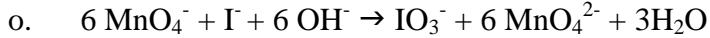
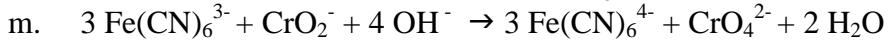
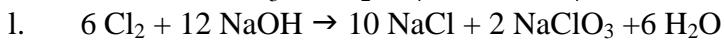
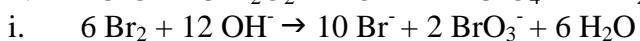
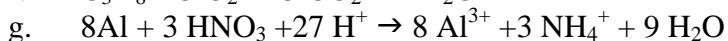
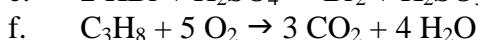
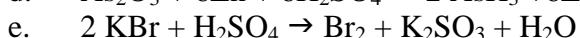
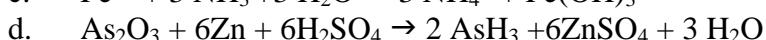
2.8 Correggere nel testo la quantita di HNO<sub>3</sub> è di **19.0 g** invece che di 9.00 g. R = 90.1%

2.9 Correggere nel testo la formula FeS<sub>2</sub> con Fe<sub>2</sub>S<sub>3</sub> R = % di Fe<sub>2</sub>S<sub>3</sub>= 93.96 %; % di FeS = 6.04 %

2.9 R = Sn 22.70%; Pb 77.30%

2.10 R = 28.4%

2.11 Reazioni bilanciate:



2.12 R = 51.9 g

2.13 R = 1.53 g

2.14 R = 22.8 g