

Esercitazioni Numeriche e di Laboratorio di Chimica di Base

Soluzioni Esercizi Capitolo 8

- 8.1 $R = 1.06$
- 8.2 $R_1 = +0.627 \text{ V}$; $R_2 = 1.03$; $R_3 = \text{SO}_4^{2-}/\text{H}_2\text{SO}_3//\text{Ag}^+/\text{Ag}$
- 8.3 $R_1 = +1.74 \text{ V}$; $R_2 = 3.98 \text{ KJ}$
- 8.4 $R = - 0.485$
- 8.5 $R_1 = + 0.602 \text{ V}$; $R_2 = \text{Cu}^{2+} + \text{Ni} \rightarrow \text{Cu} + \text{Ni}^{2+}$
- 8.6 $R_1 = \text{Ag}/\text{Ag}^+ // \text{Cu}^{2+}/\text{Cu}$; $R_2 = + 0.289 \text{ V}$
- 8.7 $R = - 0.230 \text{ V}$
- 8.8 $R = + 1.48 \text{ V}$
- 8.9 $R = 8.82 \cdot 10^{-23} \text{ M}$
- 8.10 $R_1 = + 2.26 \text{ V}$; $R_2 = \text{Zn} + \text{Cl}_2 \rightarrow \text{Zn}^{2+} + 2 \text{Cl}^-$
- 8.11 $R_1 = \text{rame solido}$; $R_2 = \text{Cu}/\text{Cu}^{2+} // \text{Cr}_2\text{O}_7^{2-}, \text{Cr}^{3+}/\text{Cr}$
 $R_3 = 3 \text{ Cu} + \text{Cr}_2\text{O}_7^{2-} + 14 \text{ H}^+ \rightarrow 2\text{Cr}^{3+} + 3 \text{ Cu}^{2+} + 7 \text{ H}_2\text{O}$
 $R_4 = +0.776 \text{ V}$
- 8.12 $R = 0.736 \text{ mol}$
- 8.13 $R_1 = 1.23 \text{ mol}$; $R_2 = 11.90$
- 8.14 $R = 1.80 \text{ g}$
- 8.15 $R = 329 \text{ min}$