

## Esercitazioni Numeriche e di Laboratorio di Chimica di Base

### Soluzioni Esercizi Capitolo 8

8.1  $R = 1.06$

8.2  $R_1 = +0.627 \text{ V} ; R_2 = 1.03 ; R_3 = \text{SO}_4^{2-}/\text{H}_2\text{SO}_3//\text{Ag}^+/\text{Ag}$

8.3  $R_1 = +1.74 \text{ V} ; R_2 = 3.98 \text{ KJ}$

8.4  $R = -0.485$

8.5  $R_1 = +0.602 \text{ V} ; R_2 = \text{Cu}^{2+} + \text{Ni} \rightarrow \text{Cu} + \text{Ni}^{2+}$

8.6  $R_1 = \text{Ag}/\text{Ag}^+ // \text{Cu}^{2+}/\text{Cu} ; R_2 = +0.289 \text{ V}$

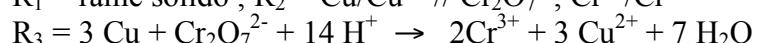
8.7  $R = -0.230 \text{ V}$

8.8  $R = +1.48 \text{ V}$

8.9  $R = 8.82 \cdot 10^{-23} \text{ M}$

8.10  $R_1 = +2.26 \text{ V} ; R_2 = \text{Zn} + \text{Cl}_2 \rightarrow \text{Zn}^{2+} + 2 \text{Cl}^-$

8.11  $R_1 = \text{rame solido} ; R_2 = \text{Cu}/\text{Cu}^{2+} // \text{Cr}_2\text{O}_7^{2-}, \text{Cr}^{3+}/\text{Cr}$



$$R_4 = +0.776 \text{ V}$$

8.12  $R = 0.736 \text{ mol}$

8.13  $R_1 = 1.23 \text{ mol} ; R_2 = 11.90$

8.14  $R = 1.80 \text{ g}$

8.15  $R = 329 \text{ min}$